

A THEORETICAL ANALYSIS OF THE REACTION BETWEEN ETHYL AND MOLECULAR OXYGEN

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Using a combination of electronic-structure theory, variational transition-state theory, and solutions to the time-dependent master equation, we have studied the kinetics of the reaction between ethyl and molecular oxygen theoretically over wide ranges of temperature and pressure. The agreement between theory and experiment is quite good. By comparing the theoretical and experimental results describing the kinetic behavior, we have been able to deduce a value for the $\text{C}_2\text{H}_5\text{--O}_2$ bond energy of ~ 34 kcal/mol and a value for the exit-channel transition-state energy of -4.3 kcal/mol (measured from reactants). These numbers compare favorably with our electronic-structure theory predictions of 33.9 kcal/mol and -3.0 kcal/mole, respectively. The master-equation solutions show three distinct temperature regimes for the reaction, discussed extensively in this paper. Above $T \approx 700$ K, the reaction can be written as an elementary step, $\text{C}_2\text{H}_5 + \text{O}_2 \leftrightarrow \text{C}_2\text{H}_4 + \text{HO}_2$, with the rate coefficient $k(T) = 3.19 \times 10^{-17} T^{1.02} \exp(2035/RT)$ cm^3 (molecules s) independent of pressure, even though the intermediate collision complex may suffer a large number of collisions.

Introduction

The reactions of alkyl radicals with molecular oxygen constitute an important class of reactions in combustion. It is the properties of these reactions that cause the change in oxidation mechanism of alkanes between 500 K and 800 K and the “negative temperature coefficient” of their overall oxidation rate in the same temperature regime [1–3]. Ethyl (C_2H_5) is the prototype alkyl radical, and consequently its reaction with O_2 has been studied extensively in the laboratory [4–12]. There have also been several attempts to treat the reaction theoretically [8,13,14], the most extensive of which is the study of Wagner et al. [8]. The latter authors provided a comprehensive review of the literature through 1990, so we may dispense with that here.

The present investigation is theoretical. It is a rather substantial extension of the work of Wagner et al., who treated the $\text{C}_2\text{H}_5 + \text{O}_2$ reaction as the sum of four separate elementary reactions, each with its own pressure- and temperature-dependent rate coefficients. These rate coefficients were calculated from chemically activated Rice–Ramsperger–Kassel–Marcus (RRKM) theory using canonical variational transition-state theory for the loose transition state and a pseudo strong-collider model to calculate

the effects of collisions. In the present investigation, we use *ab initio* electronic structure theory (density functional theory [DFT] and a G2-like method) to characterize various features of the potential energy surface, and we use microcanonical (μVT) and microcanonical/J-conservative ($\mu\text{VT-J}$) variational transition-state (RRKM) theory to calculate the high-pressure limit of the addition reaction and the zero-pressure limit of the bimolecular channel. Most importantly, however, we use solutions to the time-dependent master equation to characterize the reaction over wide ranges of temperature and pressure, including the “transition region” between 500 K and 800 K.

Theory

Figure 1 shows a potential energy diagram for the $\text{C}_2\text{H}_5 + \text{O}_2$ reaction. The energies of the stationary points on the potential energy surface (PES) come from the work described below. Our objective is to calculate the total rate coefficient for the reaction $k(T,p)$, and the branching fraction $\alpha(T,p)$, that is, the fraction that forms the bimolecular products $\text{C}_2\text{H}_4 + \text{HO}_2$.

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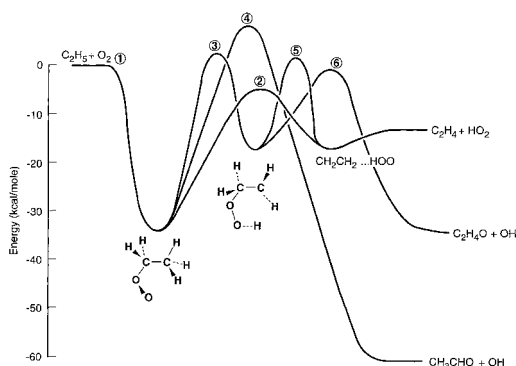


FIG. 1. Reaction coordinate diagram for the $\text{C}_2\text{H}_5 + \text{O}_2$ reaction, based primarily on the G2-like energies of Table 1. The energies shown for the stationary points of the potential energy surface are potential energy plus zero-point vibrational energy.

TABLE 1
Energetics relative to reactants for the stationary points in the $\text{C}_2\text{H}_5 + \text{O}_2$ reaction (energies in kcal/mol)

Species	B3LYP	G2 ^c -like	$\langle S^2 \rangle^d$	exp ^e
$\text{C}_2\text{H}_5\text{O}_2(\text{g})$	-29.0 ^a	-33.9	0.763	-34.0
$\text{C}_2\text{H}_5\text{O}_2(\text{t})$	-29.0 ^a	-33.8	0.763	-34.0
TS-2	-1.9 ^a	-3.0	0.855	-4.3
TS-3	8.0 ^a	3.1	0.813	
$\text{CH}_2\text{CH}_2\text{OOH}$	-9.9 ^a	-17.0	0.763	
TS-5	1.5 ^a	1.9	1.023	
$\text{CH}_2\text{CH}_2 \cdots \text{HOO}$	-14.0 ^a	-16.6	0.762	
$\text{C}_2\text{H}_4 + \text{HO}_2$	-11.2 ^a	-12.7	0.763	
TS-6	1.5 ^b	-0.6	1.297	
$\text{C}_2\text{H}_4\text{O} + \text{OH}$	-25.2 ^b	-34.0	0.757	
TS-4	14.8 ^b	8.2	0.950	
$\text{CH}_3\text{CHO} + \text{OH}$	-56.5 ^b	-61.3	0.757	
Abstraction	9.5 ^b	18.2	1.687	

^aFrom Ref. [16] using the TZ2Pf basis function.

^bPresent work employing the 6-311 + $G(d,p)$ basis set.

^cApproximate $E[\text{QCISD}(T)/6-311 + G(3df, 2pd)]$ energy as described in equation 1.

^dExpectation value of the total electron spin squared for the MP2 wave function used in the G2-like calculations.

^eDeduced from comparing predicted kinetic behavior with experiment.

Quantum Chemistry

The $\text{C}_2\text{H}_5 + \text{O}_2$ reaction has been the subject of a number of *ab initio* quantum chemical studies (cf. [15–17] and references cited therein) with methods ranging from density functional theory to more traditional MP_n , coupled cluster, and configuration interaction methods. The vibrational frequencies obtained in the B3LYP [18]/DZP analysis of Ref. [16]

are employed in the present investigation. In the interest of further delineating the allowable energetics, we have evaluated stationary-point electronic energies with an approach that is closely analogous to the G2 method [19]:

$$\begin{aligned}
 &E[\text{QCISD}(T)/6-311 + G(3df, 2pd)] \\
 &\cong E[\text{QCISD}(T)/6-311 + G(d,p)] \\
 &+ E[\text{MP2}/6-311 + G(3df, 2pd)] \\
 &- E[\text{MP2}/6-311 + G(d,p)] \\
 &+ \text{HLC} + E^{(0)}
 \end{aligned} \quad (1)$$

where $E^{(0)}$ is the zero-point vibrational energy.

These evaluations and the rate coefficient calculations discussed below employ molecular structures obtained from B3LYP/6-311 + $G(d,p)$ optimizations. The higher-level corrections (HLCs) used in different variants of G2 theory [19–21] yield a decrease in the energy (relative to reactants) of about 2.6 to 3.1 kcal/mol for each of the stationary points of interest here. In contrast, the more recent G3 theory [22] suggests the HLCs should be only -0.3 kcal/mol, and that this is likely to be a better correction even for calculations at the G2 level. Thus, this smaller correction is employed here.

The results of these G2-like calculations are provided in Table 1 and are seen to be in reasonable agreement with the corresponding B3LYP calculations of Ref. [16], but they do indicate a general lowering of the energies. Also shown in Table 1 are the stationary-point energies deduced by comparing our rate-coefficient calculations with experimental results (discussed below). These energies are in good agreement with the *ab initio* results.

The Loose Transition State

There are no electronic structure calculations from which to draw information about the PES along the $\text{C}_2\text{H}_5 + \text{O}_2$ entrance channel, and it is beyond the scope of the present work to provide such information. Consequently, we approximated the potential in this region of the surface by the method described by Miller and Klippenstein [23]. Briefly, the potential consists of three parts.

1. *The potential along the reaction coordinate.* This part of the potential is approximated by the Varshni function [23,24]. The two parameters in the Varshni potential come from knowledge of the bond energy and from a DFT calculation of the force constant matrix at the potential minimum.

2. *The potential for the “conserved” degrees of freedom orthogonal to the reaction coordinate.* This part of the potential corresponds to the degrees of freedom that can be identified as normal-mode vibrations in the separated fragments and is assumed to be the same as in the fragments.

3. *The potential for the “transitional” degrees of freedom orthogonal to the reaction coordinate.* This part of the potential is described in terms of a set of internal angles. The potential is written essentially as a sum of products (in pairs) of sinusoidal functions, with the appropriate phases and periods determined by symmetry (see equation 4 of ref. 23). The coefficients in the expression are functions of the bond distance R and are determined by calculating the appropriate force constant matrix using DFT at the potential minimum, $F_{ij}(R_0)$ and assuming that these matrix elements decay exponentially with bond distance:

$$F_{ij}(R) = F_{ij}(R_0) \exp[-\eta(R - R_0)] \quad (2)$$

The only adjustable constant is η , and we adjust its value, $\eta = 1.45 \text{ \AA}^{-1}$, to give an accurate representation of the high-pressure (p), low-temperature (T) rate coefficients.

The sum-of-states functions $N_{\Gamma}^{\pm}(E, J)$ and $N_{\Gamma}^{\pm}(E)$ are calculated with the VARIFLEX [25] computer code using methods developed by Klippenstein [26]. In the former case, there is one transition-state dividing surface, TS-1, for every E, J combination, and in the latter there is one for every E , where E is the total energy and J is the total angular momentum quantum number.

Hindered Rotors, Symmetry Numbers, and State Counting

The $\text{CH}_3 \cdots \text{CH}_2$ torsional mode in the ethyl radical is treated in the present investigation as a hindered rotor using a Pitzer-Gwinn approximation [27] for the canonical partition function of the reactants. The corresponding contribution to the transition-state partition function has been evaluated with a microcanonical analog. In particular, the hindered-rotor density of states was evaluated as the classical hindered rotor density times the quantum harmonic density divided by the classical harmonic density. The hindered-rotor rotational constant (15.2 cm^{-1}) was evaluated according to the prescription provided by Pitzer in equation 1 of Ref. [28]. A sinusoidally hindered potential of period $2\pi/3$ was employed with the barrier height set to yield the *ab initio* harmonic vibrational frequency at the minimum of the hindering potential. Rotational symmetry numbers of 3 and 2 were employed for the ethyl and oxygen fragments, respectively. The transition-state evaluations for the entrance channel included the contribution from the binding of only one of the two oxygens. Thus, the net symmetry number for the entrance channel transition state was reduced from 6 to 3. The reaction was assumed to occur only on the ground doublet electronic PES, and so an electronic correction factor of 1/3 was applied to the total rate coefficient.

For the complex $(\text{CH}_3\text{CH}_2\text{O}_2)$, we have treated both the $\text{OO} \cdots \text{CH}_2\text{CH}_3$ and the $\text{CH}_3 \cdots \text{CH}_2\text{OO}$ torsional modes as hindered rotors using a microcanonical analog of the Pitzer-Gwinn approximation. The internal rotational constants for these two hindered rotors are 2.22 cm^{-1} and 5.65 cm^{-1} , respectively. Hindering potentials of period $2\pi/3$ were employed for both these modes, since the trans and two gauche configurations are accidentally degenerate (to within 0.1 kcal/mol). The combination of this accidental degeneracy and the enantiomeric relation of the two gauche configurations yields an effective rotational symmetry number of 3. The density of states of the complex $\rho_c(E, J)$ was calculated by convolving the exact-count harmonic contributions with both the internal and external rotations. When needed, $\rho_c(E)$ was calculated as $\sum_J (2J + 1) \rho_c(E, J)$.

For the exit channel transition states (TS-2, TS-3, and TS-4), all vibrational modes were treated as harmonic oscillators, and the rotational symmetry number was unity. The sums of states $N_2^{\pm}(E, J)$, $N_3^{\pm}(E, J)$, and $N_4^{\pm}(E, J)$ were computed from conventional transition-state theory by the usual methods [29]. The corresponding $N^{\pm}(E)$ values were calculated from $N^{\pm}(E) = \sum_J (2J + 1) N^{\pm}(E, J)$.

Rate Coefficient Calculations

First, consider the two limits, the high-pressure limit of the addition reaction and the zero-pressure limit, which of course includes only the bimolecular product channel. Using $\mu\text{VT-J}$, we can write the corresponding rate coefficient expressions as

$$k_{\infty}(T) = \frac{1}{hQ_{\text{R}}(T)} \int_0^{\infty} \left[\sum_J (2J + 1) N_{\Gamma}^{\pm}(E, J) \right] \times \exp(-\beta E) dE \quad (3)$$

and

$$k_0(T) = \frac{1}{hQ_{\text{R}}(T)} \sum_J (2J + 1) \times \int_0^{\infty} \frac{N_{\Gamma}^{\pm}(E, J) N_2^{\pm}(E, J)}{N_{\Gamma}^{\pm}(E, J) + N_2^{\pm}(E, J)} \exp(-\beta E) dE \quad (4)$$

where $\beta = (k_{\text{B}}T)^{-1}$, k_{B} is Boltzmann's constant, h is Planck's constant, and Q_{R} is the reactant partition function, including relative translational contributions. In these expressions, the electronic degeneracies and symmetry numbers have been subsumed into Q_{R} and the $N^{\pm}(E, J)$'s. The μVT approximations to $k_0(T)$ and $k_{\infty}(T)$ are similar to the above except that there is no sum over J , the N^{\pm} 's are only functions of E , and there is only one TS-1 for every E . Equation 4 assumes that TS-2 gives the only contribution to the bimolecular channel and that, once this TS is traversed, a complex inevitably goes on to form $\text{C}_2\text{H}_4 + \text{HO}_2$. The latter is a safe assumption, and

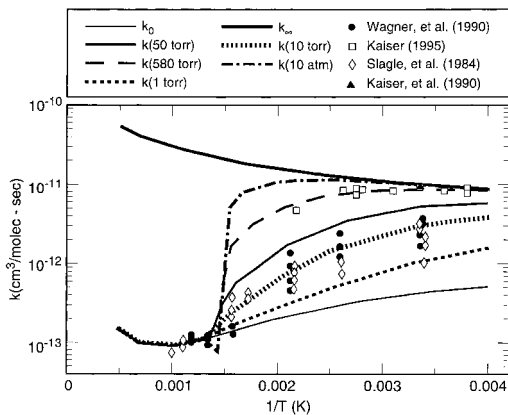


FIG. 2. Rate coefficients as a function of temperature for various pressures.

to check the former we have performed a few calculations with $N_2^\pm(E, J)$ in equation 4 replaced by $N_2^\pm(E, J) + N_3^\pm(E, J) + N_4^\pm(E, J)$. This gives us an estimate of the upper limit of the error incurred by neglecting TS-3 and TS-4 as possible exit paths from the initial complex.

To calculate rate coefficients as a function of T and p , we solve the time-dependent, one-dimensional master equation. To formulate the problem, envision the situation typical in kinetics experiments, $n_{C_2H_5} \ll n_{O_2} \ll n_{He}$ where n denotes number density. For the present investigation we assumed that the bath gas is always helium and that the reactants are constantly maintained in a thermal distribution. Under these conditions we can write a linear master equation for $n(E)$, where $n(E)dE$ is the number density of $C_2H_5O_2$ complexes between energies E and $E + dE$, and a linear rate equation for $n_{C_2H_5}$. Space limitations preclude giving a derivation here, but it is possible to express these equations in a simple form. After approximating the collision integral in the master equation as a sum, the problem can be posed in Dirac notation as

$$\frac{d}{dt} |y(t)\rangle = G|y(t)\rangle \quad (5)$$

In equation 5, $|y(t)\rangle$ is the vector of unknowns, $|y(t)\rangle \rightarrow [y_1, \dots, y_i, \dots, y_M, (K_{eq}n_{O_2}/\Delta E)^{1/2}x_{C_2H_5}]^T$, where $y_i = y(E_i)$, $E_i = E_0 + (i - 1)\Delta E$, $i = 1, \dots, M$, M is the number of grid points in the energy space, ΔE is the spacing between grid points, E_0 is the energy corresponding to the ground vibrational state of $C_2H_5O_2$, $x_{C_2H_5} = n_{C_2H_5}(t)/n_{C_2H_5}(0)$, $y(E) = x(E)/\sqrt{F(E)}$, $x(E) = n(E)/n_{C_2H_5}(0)$, $F(E)$ is the equilibrium energy distribution of the $C_2H_5O_2$ adduct, and K_{eq} is the equilibrium constant for the addition reaction, $C_2H_5 + O_2 \leftrightarrow C_2H_5O_2$. The operator G in the discrete energy basis is a real symmetric (Hermitian) matrix, which allows a simple solution.

One may write the solution to equation 5 as

$$|y(t)\rangle = \sum_{i=1}^{M+1} e^{\lambda_i t} |g_i\rangle \langle g_i | y(0) \rangle \quad (6)$$

where λ_i and $|g_i\rangle$ are the $M + 1$ (negative) eigenvalues and eigenvectors of G , and $|y(0)\rangle$ corresponds to $x_{C_2H_5}(0) = 1$. Methods for diagonalizing real, symmetric matrices are highly developed, and good software is relatively easy to find. We used the DSYEV routine from LAPACK [31] to diagonalize G .

From the solution vector $|y(t)\rangle$, one obtains

$$x_{C_2H_5}(t) = (K_{eq}n_{O_2}/\Delta E)^{-1/2} y_{M+1}(t)$$

$$x_j(t) = \sqrt{F(E_j)} y_j(t), \quad j = 1, \dots, M \quad (7)$$

$$x_{C_2H_5O_2}(t) = \sum_{i=1}^M x_i(t) \Delta E$$

and

$$x_p(t) = 1 - x_{C_2H_5}(t) - x_{C_2H_5O_2}(t)$$

where

$$x_p = \frac{n_{C_2H_4}(t)}{n_{C_2H_5}(0)} = \frac{n_{HO_2}(t)}{n_{C_2H_5}(0)}$$

In interpreting our results, we calculate $k(T, p, t)$, which is defined as

$$k(T, p, t) = -\frac{1}{n_{O_2}x_{C_2H_5}(t)} \frac{dx_{C_2H_5}(t)}{dt} \quad (8)$$

If all the C_2H_5 ultimately reacts (it does in the cases considered here), and there are good exponential decays, $k(T, p, t) = k(T, p)$ (the total rate coefficient) is constant in time. At sufficiently high pressure, in the temperature range roughly between 575 K and 700 K, multiple exponential decays occur. In these cases we characterize the rate using equation 8 at the time when $x_{C_2H_5} = 0.5$. When there is any ambiguity in the branching fraction, we define $\alpha(T, p)$, the fraction of the reaction that produces $C_2H_4 + HO_2$, as $x_p(t)$ at a time when $x_{C_2H_5} \approx 0.01$.

In all our master equation calculations, we used a standard exponential-down model for the energy transfer function with $\langle \Delta E_d \rangle = 200 \text{ cm}^{-1}$. All the calculations were done with VARIFLEX.

Results and Discussion

First, we consider the behavior of the reaction from a broad perspective. In Fig. 2, $k(T)$ is plotted for various pressures, along with a number of experimental results [6,8,10,11] at pressures ranging from slightly less than 1 torr to almost an atmosphere. The upper and lower solid curves represent $k_\infty(T)$ and

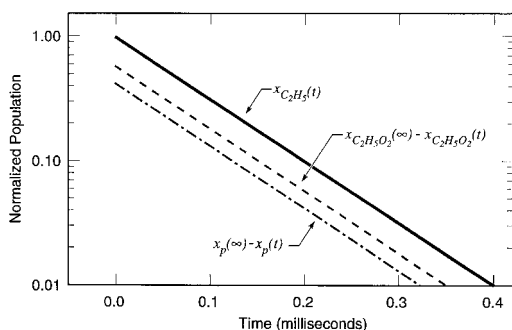


FIG. 3. Plot of $x_{C_2H_5}(t)$, $x_{C_2H_5O_2}(\infty) - x_{C_2H_5O_2}(t)$, and $x_p(\infty) - x_p(t)$ on a semi-log plot for $T = 300$ K, $p = 1$ torr. Straight lines indicate "normal" behavior.

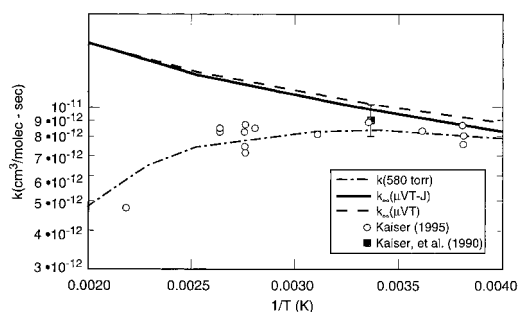


FIG. 4. Comparison of high-pressure rate coefficients with experiment.

$k_0(T)$ respectively, calculated from μVT -J. As expected, except for a pathological calculation at 700 K and 10 atm (where it was necessary to use the somewhat arbitrary definition of $k(T,p)$ discussed above), all the rate coefficients lie between these two boundaries (within reasonable uncertainty for the experiments). From the master equation calculations it is possible to identify three different regimes of the reaction.

1. *Low-temperature regime.* In this regime the reaction exhibits normal behavior. The rate coefficient and branching fraction are functions of T and p , and the time evolution is perfectly exponential (i.e., the rate coefficient is well defined). This regime persists up to $T \approx 600$ K.

2. *Transition regime.* Roughly between 600 K and 700 K the $x_{C_2H_5}(t)$ is decidedly non-exponential, at least at sufficiently high pressure. Consequently, it is impossible to extract a rate coefficient rigorously from the calculation, at least in the normal sense. The temperatures at which this occurs correspond to the equilibration of the hypothetical $C_2H_5 + O_2 \leftrightarrow C_2H_5O_2$ reaction. For the conditions of our calculations, the equilibrium fraction of C_2H_5 tied up

as $C_2H_5O_2$ at 500 K is 0.997, but it drops off precipitously to 0.108 at 700 K and to 0.0051 at 800 K. The complex behavior of the reaction in this regime is due to the simultaneous dissociation of $C_2H_5O_2$ to both $C_2H_5 + O_2$ and $C_2H_4 + HO_2$ after varying degrees of stabilization. Note the dramatic drop-off of the rate coefficient in Fig. 2 at high pressures between 650 K and 700 K as the stabilization reaction equilibrates. Stabilization becomes problematic under these conditions, and the only outlet is the bimolecular channel. The temperature range for this regime is sensitive to the C_2H_5 - O_2 bond energy. We chose the value of 34 kcal/mol used in the calculations to agree with experimental observations. It is reassuring that our G2-like calculation of 33.9 kcal/mol is in excellent agreement with this number. Previous estimates [8,30] are also compatible with these results, but the DFT bond energy of 29 kcal/mol is clearly too small.

3. *High-temperature regime.* The onset of this regime is at $T \approx 700$ K. By 750 K, the only products of the reaction are $C_2H_4 + HO_2$, and the rate-coefficient curves $k(T,p)$ in Fig. 2 for all pressures have coalesced to the zero-pressure limit. This latter result is somewhat surprising, particularly when one realizes that the $C_2H_5O_2$ complex still suffers numerous collisions at high pressure. In this regime, the reaction can be characterized by a single set of products, $C_2H_4 + HO_2$, and a single rate coefficient, $k(T)$, independent of pressure. As the last sentence implies, one again obtains good exponential decays in $x_{C_2H_5}(t)$.

Figure 3 demonstrates what we mean by normal behavior. It is a plot at $T = 300$ K, $p = 1$ torr of $x_{C_2H_5}(t)$, $x_{C_2H_5O_2}(\infty) - x_{C_2H_5O_2}(t)$, and $x_p(\infty) - x_p(t)$ on a log scale. If the reaction is governed by a single time constant (i.e., a single eigenvector of G), and one can thus identify a "good" rate coefficient, all these curves should be straight lines. They are, at least on the time scale for C_2H_5 removal.

Figure 4 is a comparison of our high- p , low- T rate coefficients with the experimental results of Kaiser and coworkers [9,11]. The agreement between theory and experiment is excellent. At 580 torr, both the experimental and theoretical rate coefficients remain relatively constant from 250 K up to $T \approx 400$ K, and then they drop off more rapidly with T . This constancy of $k(T, 580 \text{ torr})$ with T , even with a weak-collider bath gas such as helium, is a strong indication that $k_\infty(T)$ rises with temperature, an unusual result (although not unheard of) for a barrierless reaction. Our μVT -J prediction of $k_\infty(T)$ falls within the error bars of the room temperature determination of this rate coefficient reported in Ref. [9]. Wagner et al. [8] give $k_\infty(T) \approx 7 \times 10^{-12} \text{ cm}^3/(\text{molecule} \cdot \text{s})$ almost independent of temperature, a value somewhat smaller than the experimental results and the present theoretical predictions. It is also worth noting that the present μVT -J and μVT predictions of

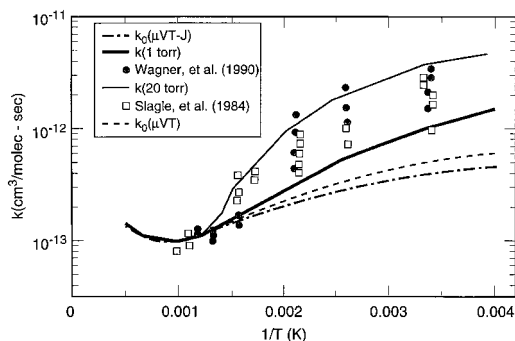


FIG. 5. Comparison of low-pressure rate coefficients with experiment.

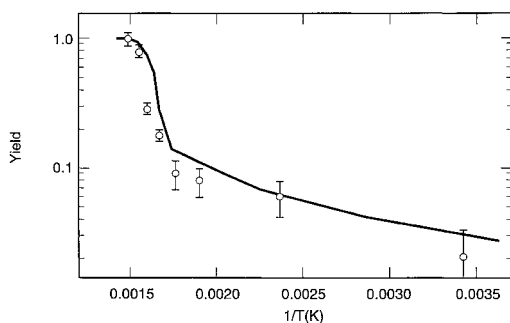


FIG. 6. Comparison of the branching fraction $\alpha(T)$ with the experiments of Clifford et al. [12] for a fixed helium number density of $1.1 \times 10^{18}/\text{cm}^3$.

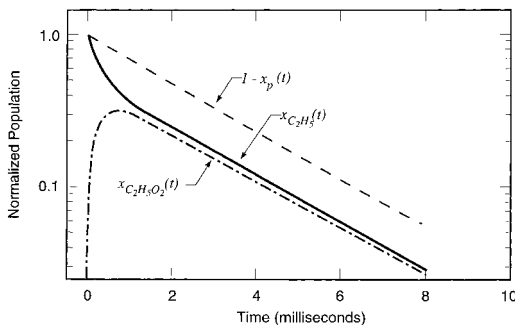


FIG. 7. Time history of $x_{\text{C}_2\text{H}_5}(t)$, $x_{\text{C}_2\text{H}_5\text{O}_2}(t)$, and $1 - x_p(t)$ for $T = 700 \text{ K}$, $p = 100 \text{ torr}$.

$k_\infty(T)$ do not differ significantly—the μVT rate coefficient is only 6% larger at 250 K and imperceptibly larger for $T > 400 \text{ K}$. The present prediction for $k_\infty(T)$ can be expressed reasonably accurately in the modified Arrhenius form, $k_\infty(T) = 4.84 \times 10^{-13} T^{0.52} \text{ cm}^3/(\text{molecule}\cdot\text{s})$ for $250 \text{ K} \leq T \leq 2000 \text{ K}$.

Figure 5 is a comparison of our theoretical predictions with the low-pressure rate coefficients measured by Gutman and collaborators [6–8]. All their

experiments were conducted at pressures ranging from slightly less than 1 torr to approximately 15 torr. Our theoretical curves corresponding to $p = 1 \text{ torr}$ and $p = 20 \text{ torr}$ should more or less bracket these results. Within reasonable experimental uncertainty, they do. The rate coefficients in the high-temperature regime ($T > 700 \text{ K}$) are governed almost exclusively by TS-2, and we adjusted the value of its energy, E_2 , in our calculations to give good agreement with the experimental rate coefficients in this regime. This value of $E_2 = -4.3 \text{ kcal/mol}$ (measured from reactants) is in reasonably good agreement with our G2-like *ab initio* result of $E_2 = -3.0 \text{ kcal}$, but it differs somewhat from the DFT calculation of -1.9 kcal/mol . Using $E_2 = -1.9 \text{ kcal/mole}$ in the calculations gives rate coefficients in the high-temperature regime that are substantially smaller than those shown in the figure (by approximately a factor of 4 at $T = 1000 \text{ K}$).

Also shown in Fig. 5 are the μVT and $\mu\text{VT-J}$ results for $k_0(T)$. The μVT rate coefficient is 29% larger than that calculated from $\mu\text{VT-J}$ at $T = 250 \text{ K}$, a difference that diminishes with increased temperature until it is virtually imperceptible for $T > 500 \text{ K}$. The two rate coefficients come together at high temperature because $k_0(T)$ begins to be controlled exclusively by the single tight transition state TS-2, that is the flux through TS-1 becomes effectively infinite (see equation 4), and we are left with a conventional transition-state theory result, with TS-2 the transition state. The contribution to $k_0(T)$ from TS-3 and TS-4 is no more than 9% at 2000 K (calculated using the G2-like barrier heights of Table 1); it is much less at lower temperatures. For $700 \text{ K} < T < 2000 \text{ K}$, we can express $k_0(T)$ as $k_0(T) = 3.19 \times 10^{-17} T^{1.02} \exp(2035/RT) \text{ cm}^3/(\text{molecule}\cdot\text{s})$, which should be sufficiently accurate for flame modeling.

The transition-state energy E_2 also plays a major role in determining the branching fraction $\alpha(T, p)$ in the low-temperature regime. Fig. 6 compares our predictions for α with the experimental results of Clifford et al. [12], who measured the branching fraction from room temperature to 700 K at a constant helium number density of $1.1 \times 10^{18}/\text{cm}^3$. The agreement between theory and experiment is quite satisfactory. Both show a slow rise in α from room temperature to about 575 K, followed by the very rapid rise between 575 K and 700 K that occurs as the $\text{C}_2\text{H}_5 + \text{O}_2 \leftrightarrow \text{C}_2\text{H}_5\text{O}_2$ reaction equilibrates.

Figure 7 shows the time history of $x_{\text{C}_2\text{H}_5}(t)$, $x_{\text{C}_2\text{H}_5\text{O}_2}(t)$, and $1 - x_p(t)$ on a semi-log plot for $T = 700 \text{ K}$, $p = 100 \text{ torr}$. This plot is characteristic of the behavior we observe throughout the transition regime (at sufficiently high pressure). The reactant concentration shows a distinctly non-exponential decay. The function $x_{\text{C}_2\text{H}_5}(t)$ first decays rapidly as $x_{\text{C}_2\text{H}_5\text{O}_2}(t)$ rises. After $x_{\text{C}_2\text{H}_5\text{O}_2}(t)$ peaks, both $x_{\text{C}_2\text{H}_5}(t)$ and $x_{\text{C}_2\text{H}_5\text{O}_2}(t)$ decay with the same exponential time constant. The most interesting result from Fig. 7 is that the bimolecular products, $x_p(t)$, are governed by this same long-term time constant even at very short

times. However, at lower pressures $x_p(t)$ can also display two-time-constant behavior, more pronounced in the 600 K to 650 K temperature range than at 700 K.

It is instructive to interpret our results in terms of the eigenvalues and eigenvectors of G . Two eigenvalues play a major role. At low temperature, it is the second largest eigenvalue of G (i.e., the second least negative), λ_2 , that determines the time evolution of both the reactants and products. The largest eigenvalue, λ_1 , is associated with the thermal dissociation of C₂H₅O₂ to C₂H₄ + HO₂ and is much larger than λ_2 . This largest eigenvalue decreases rapidly with increasing temperature (at least above 500 K), and in the transition regime λ_1 and λ_2 are of comparable size. We thus observe the behavior described above. In the high-temperature regime, there is again a separation of the eigenvalues, but now it is λ_1 that governs the time evolution of C₂H₅ and O₂ to C₂H₄ + HO₂.

A question about the importance of the direct abstraction channel always arises in discussions of alkyl radical + O₂ reactions. For the present case, we have used the G2-like transition state of Table 1 to calculate a direct abstraction rate coefficient

$$k_a(T) = 1.93 \times 10^{-17} T^{1.57} \exp(-20578/RT)$$

for 1000 K $\leq T \leq$ 2000 K. This rate coefficient is about a factor of 10 smaller at 2000 K than the $k(T)$ discussed above. The difference is even greater at lower temperatures, and thus it is probably reasonable to neglect $k_a(T)$ in most applications.

Concluding Remarks

We have used a variety of theoretical techniques to study the reaction between ethyl and O₂ over very wide ranges of temperature and pressure. These include electronic-structure theory, variational transition-state theory, and solutions to the time-dependent master equation. Our predictions of rate coefficients and product distributions are in very good agreement with experiment. Moreover, the values of the C₂H₅–O₂ bond energy and the exit-channel transition-state energy of 34 kcal/mol and –4.3 kcal/mol (measured from reactants), respectively, deduced from the master equation results, compare favorably with the values of 33.9 kcal/mol and –3.0 kcal/mol calculated from our G2-like method.

The most interesting aspect of our analysis is the prediction of three different regimes of the reaction, discussed extensively in the text. For temperatures larger than $T \approx 700$ K, the reaction can be written as a simple elementary step, C₂H₅ + O₂ \leftrightarrow C₂H₄ + HO₂, with a rate coefficient $k(T)$ independent of pressure, even though the reaction goes through an intermediate complex that may suffer numerous collisions. Moreover, the rate coefficient is the same as

that at the zero-pressure (collisionless) limit. In modified Arrhenius form, we can write $k(T)$ in this regime as $k(T) = 3.19 \times 10^{-17} T^{1.02} \exp(2035/RT)$ cm³/(molecule-s)

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COMMENTS

Jürgen Troe, *University of Göttingen, Germany*. I would like to remind you that the reaction $\text{C}_2\text{H}_5 + \text{O}_2$, which you describe, corresponds completely to the reaction $\text{CO} + \text{OH}$, that is, biexponential kinetics, pressure-dependent, pseudo-second-order rate coefficients, and low and high pressure limiting-ranges are observed. In my analysis of the $\text{CO} + \text{OH}$ reaction [1], I came to the conclusion that a two-dimensional treatment with separate energy (E) and angular momentum (J) dependencies of the rate constants is required, in particular if a loose entrance/rigid exit complex-forming bimolecular reaction is considered. Can your code handle this? Another question: what energy transfer parameters did you use?

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Author's Reply. The treatment you use in your paper is equivalent to what we call $\mu\text{VT-J}$ theory (microcanonical variational theory with J conservation), with a pseudo-strong-collider approximation for stabilization. We certainly *do* have such a capability. We conclude that J conservation is not a very important factor in the present reaction (see Fig. 5 of the paper). Our full one-dimensional master equation treatment goes beyond such a simple treatment, at least as long as J conservation is not a major concern.

For lack of anything better to use at this point, we assumed a single-exponential down model for the energy

transfer function with $\Delta E_d = 200 \text{ cm}^{-1}$. We arrived at this number by testing against some thermally activated dissociation/recombination reactions with helium as the bath gas. It worked well in the present calculations, so we did not alter it.

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Horst Hippler, *University of Karlsruhe, Germany*. Collisional energy transfer controls the transition between pressure-independent low- and high-pressure limiting rate constants. How has this been treated, and how sensitive is the product distribution on the collisional energy transfer parameters in this range? Secondly, at high temperatures the kinetic system converts into a reaction out of a pre-equilibrium. Therefore, the rate constant for HO_2 formation is not only controlled by the properties of TS2 but also by the equilibrium constant for $\text{C}_2\text{H}_5 + \text{O}_2 \leftrightarrow \text{C}_2\text{H}_5\text{O}_2$. Thus, also thermochemistry of $\text{C}_2\text{H}_5\text{O}_2$ equilibrium controls the reaction rate.

Author's Reply. Regarding the first part of your question, see the response to the question by Jürgen Troe. Secondly, I do not know what you mean by “preequilibrium”. As alluded to in the paper, the rate coefficient in the high-temperature regime does not depend on the thermochemistry of $\text{C}_2\text{H}_5\text{O}_2$. However, the temperature at which the high-temperature regime is reached does. Specifically, the temperature at which the transition regime occurs depends significantly on the $\text{C}_2\text{H}_5\text{O}_2$ bond energy.